

Formula Writing**Mixed Compounds****Learning Target**

Apply the rules for writing formulas of binary ionic compounds.

Practice writing the names and formulas of ionic and molecular compounds using the table below. Make sure to include the correct charges on each cation and anion for ionic compounds.

	Compound	Cation	Anion	Chemical Formula
1	Sodium hypochlorite	Na^+	ClO^-	NaClO
2	Xenon hexafluoride	XXXXXXXXXX	XXXXXXXXXX	XeF_6
3	Calcium bicarbonate	Ca^{2+}	HCO_3^-	$\text{Ca}(\text{HCO}_3)_2$
4	Oxygen difluoride	XXXXXXXXXX	XXXXXXXXXX	OF_2
5	Silicon dioxide	XXXXXXXXXX	XXXXXXXXXX	SiO_2
6	Palladium (II) selenide	Pd^{2+}	Se^{2-}	PdSe
7	Radium phosphate	Ra^{2+}	PO_4^{3-}	$\text{Ra}_3(\text{PO}_4)_2$
8	Copper (I) dichromate	Cu^+	$\text{C}_2\text{O}_7^{2-}$	$\text{Cu}_2\text{C}_2\text{O}_7$
9	Carbon tetrafluoride	XXXXXXXXXX	XXXXXXXXXX	CF_4
10	Iridium (III) sulfate	Ir^{3+}	SO_4^{2-}	$\text{Ir}_2(\text{SO}_4)_3$
11	Manganese (III) acetate	Mn^{3+}	$\text{C}_2\text{H}_3\text{O}_2^-$	$\text{Mn}(\text{C}_2\text{H}_3\text{O}_2)_3$
12	Dinitrogen monoxide	XXXXXXXXXX	XXXXXXXXXX	N_2O
13	Iron (III) sulfite	Fe^{3+}	SO_3^{2-}	$\text{Fe}_2(\text{SO}_3)_3$
14	Lead (IV) carbonate	Pb^{4+}	CO_3^{2-}	$\text{Pb}(\text{CO}_3)_2$
15	Sulfur trioxide	XXXXXXXXXX	XXXXXXXXXX	SO_3
16	Ammonium chromate	NH_4^+	CrO_4^{2-}	$(\text{NH}_4)_2\text{CrO}_4$
17	Tetraphosphorus decoxide	XXXXXXXXXX	XXXXXXXXXX	P_4O_{10}
18	Chlorine monofluoride	XXXXXXXXXX	XXXXXXXXXX	ClF
19	Disulfur decachloride	XXXXXXXXXX	XXXXXXXXXX	S_2Cl_{10}
20	Beryllium cyanide	Be^{2+}	CN^-	$\text{Be}(\text{CN})_2$