

**Naming****Binary Ionic Compounds****Learning Target**

Apply the rules for naming binary ionic compounds.

Practice naming binary ionic compounds using the table below. Make sure to include the correct charges on each cation and anion.

|    | Compound                       | Cation    | Anion     | Chemical Name             |
|----|--------------------------------|-----------|-----------|---------------------------|
| 1  | KI                             | $K^+$     | $I^-$     | Potassium iodide          |
| 2  | AlN                            | $Al^{3+}$ | $N^{3-}$  | Aluminum nitride          |
| 3  | BaO                            | $Ba^{2+}$ | $O^{2-}$  | Barium oxide              |
| 4  | CaBr <sub>2</sub>              | $Ca^{2+}$ | $Br^-$    | Calcium bromide           |
| 5  | RaS                            | $Ra^{2+}$ | $S^{2-}$  | Radium sulfide            |
| 6  | Rb <sub>2</sub> O              | $Rb^+$    | $O^{2-}$  | Rubidium oxide            |
| 7  | AlCl <sub>3</sub>              | $Al^{3+}$ | $Cl^-$    | Aluminum chloride         |
| 8  | Na <sub>3</sub> P              | $Na^+$    | $P^{3-}$  | Sodium phosphide          |
| 9  | Be <sub>3</sub> N <sub>2</sub> | $Be^{2+}$ | $N^{3-}$  | Beryllium nitride         |
| 10 | RbF                            | $Rb^+$    | $F^-$     | Rubidium fluoride         |
| 11 | PbO                            | $Pb^{2+}$ | $O^{2-}$  | Lead (II) oxide           |
| 12 | MnP                            | $Mn^{3+}$ | $P^{3-}$  | Manganese (III) phosphide |
| 13 | Li <sub>3</sub> N              | $Li^+$    | $N^{3-}$  | Lithium nitride           |
| 14 | CuSe                           | $Cu^{2+}$ | $Se^{2-}$ | Copper (II) selenide      |
| 15 | FeCl <sub>3</sub>              | $Fe^{3+}$ | $Cl^-$    | Iron (III) chloride       |
| 16 | CrF <sub>3</sub>               | $Cr^{3+}$ | $F^-$     | Chromium (III) fluoride   |
| 17 | Cu <sub>2</sub> O              | $Cu^+$    | $O^{2-}$  | Copper (I) oxide          |
| 18 | PtBr <sub>4</sub>              | $Pt^{4+}$ | $Br^-$    | Platinum (IV) bromide     |
| 19 | Pb <sub>3</sub> N <sub>4</sub> | $Pb^{4+}$ | $N^{3-}$  | Lead (IV) nitride         |
| 20 | AuI <sub>3</sub>               | $Au^{3+}$ | $I^-$     | Gold (III) iodide         |