

Balancing Chemical Reactions*Practice I***Learning Target**

Describe the steps of writing and balancing a chemical equation.

Directions: Balance the following chemical equations.Reactants:

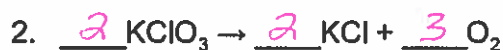
N = ~~2~~ 2

H = ~~2~~ 6

Products:

N = ~~1~~ 2

H = ~~3~~ 6

Reactants:

K = ~~1~~ 2

Cl = ~~1~~ 2

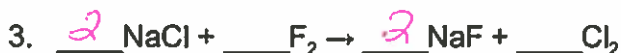
O = ~~3~~ 6

Products:

K = ~~1~~ 2

Cl = ~~1~~ 2

O = ~~2~~ 6

Reactants:

Na = ~~1~~ 2

Cl = ~~1~~ 2

F = 2

Products:

Na = ~~1~~ 2

Cl = 2

F = ~~1~~ 2

Reactants:

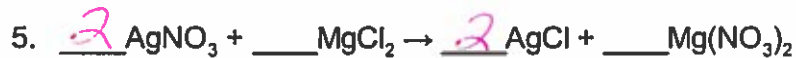
H = ~~2~~ 4

O = 2

Products:

H = ~~2~~ 4

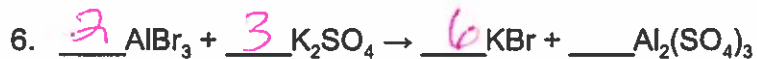
O = ~~1~~ 2

Reactants:

$\text{Cl} = 2$
 $\text{Ag} = 2$
 $\text{N} = 2$
 $\text{O} = 6$
 $\text{Mg} = 1$

Products:

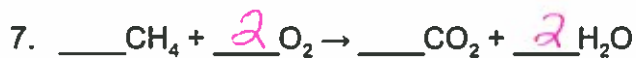
$\text{Cl} = 2$
 $\text{Ag} = 2$
 $\text{N} = 2$
 $\text{O} = 6$
 $\text{Mg} = 1$

Reactants:

$\text{Al} = 2$
 $\text{Br} = 6$
 $\text{K} = 6$
 $\text{S} = 3$
 $\text{O} = 12$

Products:

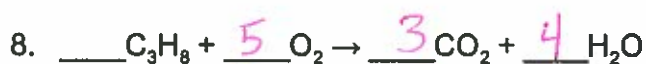
$\text{Al} = 2$
 $\text{Br} = 6$
 $\text{K} = 6$
 $\text{S} = 3$
 $\text{O} = 12$

Reactants:

$\text{C} = 1$
 $\text{H} = 4$
 $\text{O} = 4$

Products:

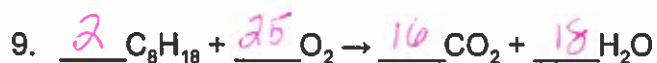
$\text{C} = 1$
 $\text{H} = 4$
 $\text{O} = 4$

Reactants:

$\text{C} = 3$
 $\text{H} = 8$
 $\text{O} = 10$

Products:

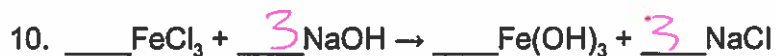
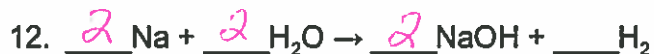
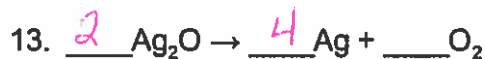
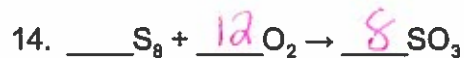
$\text{C} = 3$
 $\text{H} = 8$
 $\text{O} = 10$

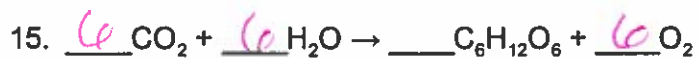
Reactants:

$\text{C} = 16$
 $\text{H} = 36$
 $\text{O} = 50$

Products:

$\text{C} = 16$
 $\text{H} = 36$
 $\text{O} = 50$

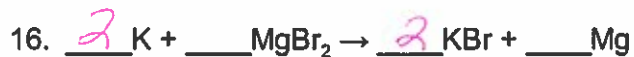
Reactants:Fe = 1
Cl = 3
Na = ~~X~~ 3
O = ~~X~~ 3
H = ~~X~~ 3Products:Fe = 1
Cl = ~~X~~ 3
Na = ~~X~~ 3
O = 3
H = 3Reactants:P = ~~X~~ 4
O = ~~X~~ 10Products:P = ~~X~~ 4
O = ~~X~~ 10Reactants:Na = ~~X~~ 2
H = ~~X~~ 4
O = ~~X~~ 2Products:Na = ~~X~~ 2
H = 3 4
O = ~~X~~ 2Reactants:Ag = ~~X~~ 4
O = ~~X~~ 2Products:Ag = ~~X~~ 4
O = 2Reactants:S = 8
O = ~~X~~ 24Products:S = ~~X~~ 8
O = ~~X~~ 24

Reactants:

$$\begin{aligned} \text{C} &= 12 \\ \text{O} &= 36 \\ \text{H} &= 12 \end{aligned}$$

Products:

$$\begin{aligned} \text{C} &= 6 \\ \text{O} &= 18 \\ \text{H} &= 12 \end{aligned}$$

Reactants:

$$\begin{aligned} \text{K} &= 2 \\ \text{Mg} &= 1 \\ \text{Br} &= 2 \end{aligned}$$

Products:

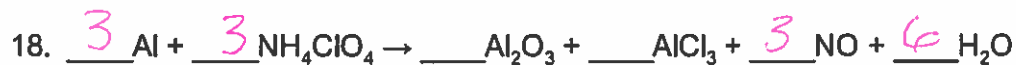
$$\begin{aligned} \text{K} &= 2 \\ \text{Mg} &= 1 \\ \text{Br} &= 2 \end{aligned}$$

Reactants:

$$\begin{aligned} \text{H} &= 2 \\ \text{Cl} &= 2 \\ \text{Ca} &= 1 \\ \text{C} &= 1 \\ \text{O} &= 3 \end{aligned}$$

Products:

$$\begin{aligned} \text{H} &= 2 \\ \text{Cl} &= 2 \\ \text{Ca} &= 1 \\ \text{C} &= 1 \\ \text{O} &= 3 \end{aligned}$$

Reactants:

$$\begin{aligned} \text{Al} &= 3 \\ \text{N} &= 3 \\ \text{H} &= 12 \\ \text{Cl} &= 3 \\ \text{O} &= 12 \end{aligned}$$

Products:

$$\begin{aligned} \text{Al} &= 3 \\ \text{N} &= 3 \\ \text{H} &= 12 \\ \text{Cl} &= 3 \\ \text{O} &= 12 \end{aligned}$$