

Balancing & Identifying Reactions**Practice II****Learning Target**

Describe the steps of writing and balancing a chemical equation.
Describe the five general types of reactions.

Directions: Balance the following equations and state the reaction type.



Ca = 1
N = 2
O = 10
Na = 2
S = 1

Ca = 1
N = ~~2~~
O = ~~10~~
Na = ~~2~~
S = 1

Reaction type: Double Replacement



Fe = ~~1~~
O = 2

Fe = ~~1~~
O = ~~2~~

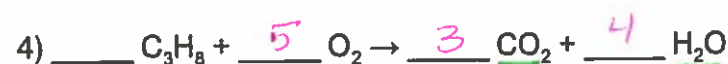
Reaction type: Synthesis



S = 1
O = 4
H = 2

S = 1
O = 4
H = 2

Reaction type: Synthesis



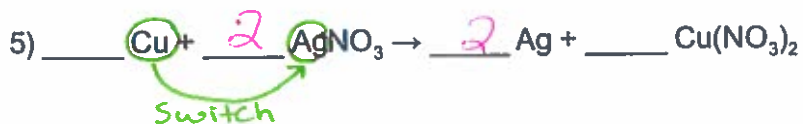
C = 3
H = 8

C = ~~1~~
H = ~~8~~

O = ~~10~~

O = ~~10~~

Reaction type: Combustion



Cu = 1
Ag = $\times 2$
N = $\times 2$
O = $\times 6$

Cu = 1
Ag = $\times 2$
N = 2
O = 6

Reaction type: Single Replacement



K = $\times 2$
Cl = $\times 2$
O = $\times 6$

K = $\times 2$
Cl = $\times 2$
O = $\times 6$

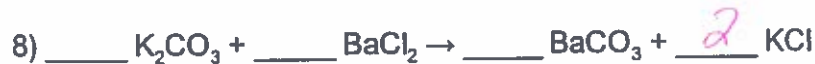
Reaction type: Decomposition



Mg = 1
H = $\times 2$
Cl = $\times 2$

Mg = 1
H = 2
Cl = 2

Reaction type: Single Replacement



K = 2
C = 1
O = 3
Ba = 1
Cl = 2

K = $\times 2$
C = 1
O = 3
Ba = 1
Cl = $\times 2$

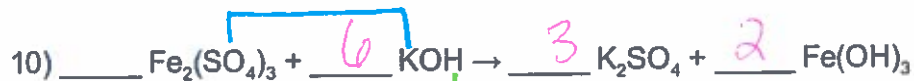
Reaction type: Double Replacement



Zn = 1
C = 1
O = 3

Zn = 1
C = 1
O = 3

Reaction type: Decomposition



Fe = 2
S = 3
O = ~~18~~
K = $\times 6$
H = $\times 6$

Fe = $\times 2$
S = $\times 3$
O = $\times 18$
K = $\times 6$
H = $\times 6$

Reaction type: Double Replacement