

**Naming****Ternary Ionic Compounds****Learning Target**

Apply the rules for writing formulas of binary ionic compounds.

Practice writing ternary ionic compounds using the table below. Make sure to include the correct charges on each cation and anion.

	Compound	Cation	Anion	Chemical Name
1	$K_2CO_3$	$K^+$	$CO_3^{2-}$	Potassium carbonate
2	KOH	$K^+$	$OH^-$	Potassium hydroxide
3	$Ca(NO_3)_2$	$Ca^{2+}$	$NO_3^-$	Calcium nitrate
4	$SrSO_4$	$Sr^{2+}$	$SO_4^{2-}$	Strontium sulfate
5	$CuHCO_3$	$Cu^+$	$HCO_3^-$	Copper (I) bicarbonate
6	$Li_3PO_4$	$Li^+$	$PO_4^{3-}$	Lithium phosphate
7	$Fe(OH)_2$	$Fe^{2+}$	$OH^-$	Iron (II) hydroxide
8	$Na_2SO_3$	$Na^+$	$SO_3^{2-}$	Sodium sulfite
9	$Sn(NO_2)_4$	$Sn^{4+}$	$NO_2^-$	Tin (IV) nitrite
10	$Ca(SCN)_2$	$Ca^{2+}$	$SCN^-$	Calcium thiocyanate
11	$Cu(CH_3COO)_2$	$Cu^{2+}$	$CH_3COO^-$	Copper (II) acetate
12	$NH_4OH$	$NH_4^+$	$OH^-$	Ammonium hydroxide
13	$Cs_2S_2O_3$	$Cs^+$	$S_2O_3^{2-}$	Cesium thiosulfate
14	$Fe_2(CO_3)_3$	$Fe^{3+}$	$CO_3^{2-}$	Iron (III) carbonate
15	$CuMnO_4$	$Cu^+$	$MnO_4^-$	Copper (I) permanganate
16	$Fe(OH)_3$	$Fe^{3+}$	$OH^-$	Iron (III) hydroxide
17	$Cu_2O$	$Cu^+$	$O^{2-}$	Copper (I) oxide
18	$Pb(CN)_2$	$Pb^{2+}$	$CN^-$	Lead (II) cyanide
19	$(NH_4)_2CO_3$	$NH_4^+$	$CO_3^{2-}$	Ammonium carbonate
20	$Rb_2Cr_2O_7$	$Rb^+$	$Cr_2O_7^{2-}$	Rubidium dichromate